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The number of collisions between reactants and catalyst is at a maximum because the catalyst is uniformly dispersed throughout the reaction mixture. Many homogeneous catalysts in industry are transition metal compounds (Table

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\\(\PageIndex{2}\\\)), but recovering these expensive catalysts from solution has been a major challenge.

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solution and 100mL of the 6% hydrogen peroxide solution. 3. Heat to the desired temperature {50°C-70°C). 4. Remove from the heat and add the catalyst

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(CoCl_2 , solution) to
one of the beakers.
Place the reaction
vessels in a shallow
container in case of
overflow. The reaction
is very vigorous at
higher temperatures
Discussion

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iterative solution of the

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suggested method is significantly close to the actual solution of the kinetic models in most cases.

Chemical Kinetics

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and B? $A + B \rightarrow E$ (fast)

$E + B \rightarrow C + D$ (slow) a.

What is the order with respect to reactants A and B? 1, 2 b.

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Catalysis -

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The conventional steady-state flow methods that are used widely to investigate the kinetics of heterogeneous catalytic reactions are limited, because under steady-state conditions, the elementary steps proceeding in series take place at the same rate, and the overall

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kinetics provides
somewhat limited
insight into the
mechanism.

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mathematicians and
chemists whose

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